

Magnetic Properties of the Dioxime Complex 2,3-Diamino-1,4-diazabuta-1,3-diene-1,4-diol-bis(2,3-diamino-4-hydroxy-1,4-diazabuta-1,3-dien-1-olato)cobalt(II)

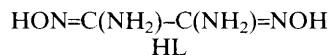
Takahiro Hirotsu,*^a Norio Takagi,^a Jitsuo Sakakibara,^a Shunsaku Katoh^a and Manabu Senō^b

^a Government Industrial Research Institute, Shikoku, 2-3-3 Hananomiya-cho, Takamatsu 761, Japan

^b Institute of Industrial Science, The University of Tokyo, 7-22-1 Roppongi, Minato-ku, Tokyo 106, Japan

Crystals of the title complex in the low-spin state ($S = 1/2$) exhibit one-dimensional ferromagnetic interactions along the b -axis.

The development of magnetic interactions in metal-ion complexes with a stack arrangement is of interest in exploring the design of organic ferromagnetic compounds. Metal-ion complexes with glyoximes are well known to possess a variety of stack structures;¹ however their magnetic properties have not been examined in detail. In this study, we report that the title compound, $[\text{CoL}_2(\text{HL})]$, with a slipped-stack arrangement is an $S = 1/2$ 1-D ferromagnet along the stacking axis.



$[\text{CoL}_2(\text{HL})]$ was obtained in 35–40% yield as dark purple crystals from a warm aqueous solution (100 cm^3) containing $[\text{Co}^{\text{II}}(\text{OAc})_2] \cdot 4\text{H}_2\text{O}$ (10 mmol) and the dioxime HL (20 mmol) under nitrogen,² and its composition was confirmed by IR spectroscopy and elemental analyses to be $[\text{Co}^{\text{II}}(\text{C}_2\text{H}_5\text{N}_4\text{O}_2)_2(\text{C}_2\text{H}_6\text{N}_4\text{O}_2)]$. The complex is monoclinic, space group $\text{C}2/c$ and its structure consists of chains of planar $\text{Co}^{\text{II}}\text{L}_2$ units stacked along the b -axis, through long Co–N bonds to terminal N atoms of the adjacent molecular units; the stacked chains form sheets in the bc plane.²

The ESR behaviour of $[\text{CoL}_2(\text{HL})]$ crystals was examined over the temperature range 123–298 K. The spectra were characterized by the combination of a broad absorption band centred at an average g value (g_{av}) of 2.35 with a very weak but distinct hyperfine splitting (hfs) pattern centred at a g_z value of 2.012, which results from the interaction of the unpaired electron with the ^{59}Co nucleus ($I = 7/2$), with $A_{\text{Co}} = 12 \text{ mT}$.

The absence of any hfs signal from ^{14}N ($I = 1$) nuclei on the apical sites means that the interaction between the Co and the N atoms is as weak as in the case of $\text{Co}^{\text{II}}(\text{Hdmg})_2$ ($\text{H}_2\text{dmg} = \text{butane-2,3-dione dioxime}$) in frozen water and alcoholic solvents and much weaker than that observed in strongly coordinating media like pyridine.³ The g_{av} value of 2.35

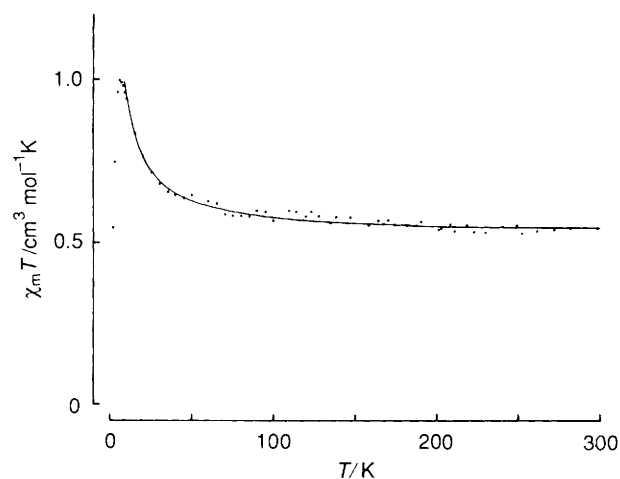


Fig. 1 Molar susceptibility \times temperature vs. temperature for $[\text{CoL}_2(\text{HL})]$. The solid line is theoretical.

remains constant over the above temperature range, while the hfs signal centred at $g_z = 2.012$ weakens significantly with a decrease in temperature. Provided that the system is isotropic, the S value can be estimated through the relationship $\mu_{\text{eff}}/\mu_B = g[S(S+1)]^{1/2}$, where μ_{eff} is the effective magnetic moment. If a μ_{eff} value of $2.05 \mu_B$, as described below, and $g = g_{\text{av}} = 2.35$, the S value is calculated to be 0.505, which is consistent with the $S = 1/2$ ground state, *i.e.*, the low-spin state.

The magnetic properties of $[\text{CoL}_2(\text{HL})]$ crystals were examined by means of a Squid magnetometer at 2–100 K and a VSM instrument at 77–298 K. $\mu_{\text{eff}} = 2.05 \mu_B$ at 298 K, consistent with the values of Co^{II} complexes with a square-planar coordination geometry. The magnetic data are plotted as $\chi_m T$ (χ_m = molar magnetic susceptibility, T = temperature) *vs.* T in Fig. 1. A marked rise in $\chi_m T$ at temperatures lower than 50 K indicates the transition to a phase with dominant ferromagnetic interactions. It is shown on the basis of crystallographic information on $[\text{CoL}_2(\text{HL})]$ that this result could be reasonably interpreted in terms of a dominant ferromagnetic spin-ordering within the stack chain along the b -axis and a secondary antiferromagnetic interaction between the chains in the bc plane. Thus, we assume tentatively the isotropic $S = 1/2$ Heisenberg model based on a high-temperature Padé expansion (eqn. 1),⁴ modified by the addition of a mean-field correction (eqn. 2) to account for the interchain interactions,⁵ where $K = J/2kT$; J is the ferromagnetic

$$\chi_m' T = (N \mu_B^2 g^2 / 4k) [(1.0 + 5.799916 K + 16.902653 K^2 + 29.376885 K^3 + 29.832959 K^4 + 14.036918 K^5) / (1.0 + 2.7979916 K + 7.0086780 K^2 + 8.6538644 K^3 + 4.5743114 K^4)]^{2/3} \quad (1)$$

$$\chi_m(\text{corr}) = \chi_m' / [1 - (2ZJ' / Ng^2 \mu_B^2) \chi_m'] \quad (2)$$

coupling constant, Z is the number of nearest-neighbour chains ($Z = 2$ in this study), J' is the exchange integral for the magnetic interaction between the nearest-neighbour chains, and the other symbols have their usual meanings. The parameter values were determined to be consistent with the observed results in Fig. 1 as follows: $J = 9.04 \text{ cm}^{-1}$ and $J' = -0.005 \text{ cm}^{-1}$ at temperatures 10–298 K with a value of $g = g_{\text{av}} = 2.35$. The good agreement indicates the existence of ferromagnetic intrachain interactions along the b -axis accompanied by a very weak interchain antiferromagnetic interaction in the bc plane. The principal mechanism of superexchange could operate through the interaction of the $\text{Co } d_{z^2}$ (the highest partially occupied magnetic orbital) with the two apical N atoms which form the π systems of the adjacent complex molecules along the stack chain, similar to the $(\text{pda})\text{Cu}(\text{ox})_2$ system (pda = propylenediammonium; ox = oxalate).⁶

Received, 30th May 1990; Com. 0/02421J

References

- 1 B. M. Hoffman, J. Martinsen, L. J. Pace and J. A. Ibers, in *Extended Linear Chain Compounds*, vol. 3, ed. J. S. Miller, Plenum Press, New York, 1983, ch. 10.
- 2 Ö. Bekaroglu, S. Sarisaban, A. R. Koray and M. L. Ziegler, *Z. Naturforsch., Teil B.*, 1977, **32**, 387.
- 3 W. Lubitz, C. J. Winscom, H. Diegruber and R. Moseler, *Z. Naturforsch., Teil A*, 1987, **42**, 970, and references cited therein.
- 4 G. A. Baker, Jr., G. S. Rushbrooke and H. E. Gilbert, *Phys. Rev. A.*, 1964, **135**, 1272.
- 5 D. D. Swank, C. P. Landee and R. D. Willett, *Phys. Rev. B*, 1979, **20**, 2154.
- 6 D. R. Bloomquist, J. J. Hansen, C. P. Landee, R. D. Willett and R. Buder, *Inorg. Chem.*, 1981, **20**, 3308.